

Lab Preparation for Science 10 Experiment
ANALYSIS OF SOLUTIONS

Solution	Preparation
0.1 M AgNO ₃	Dilute 4.25 g of silver nitrate to 250 mL
1.0 M HCl	Dilute 20.8 mL of concentrated hydrochloric acid to 250 mL
0.5 M Na ₂ CO ₃	Dilute 15.50 g of sodium carbonate monohydrate to 250 mL
0.1 M Pb(NO ₃) ₂	Dilute 8.28 g of lead nitrate to 250 mL
1.0 M NaCl	Dilute 14.63 g of sodium chloride to 250 mL
0.1 M KI	Dilute 4.15 g of potassium iodide to 250 mL

Divide 125 mL of each stock solution as follows.

- 12 dropper bottles labeled "silver nitrate"
- 12 dropper bottles labeled "hydrochloric acid"
- 12 dropper bottles labeled "sodium carbonate"
- 12 dropper bottles labeled "lead(II) nitrate"
- 12 dropper bottles labeled "sodium chloride"
- 12 dropper bottles labeled "potassium iodide"

Divide the remaining 125 mL of each stock solution as follows.

- 12 dropper bottles containing HCl and labeled "UNKNOWN A"
- 12 dropper bottles containing Pb(NO₃)₂ and labeled "UNKNOWN B"
- 12 dropper bottles containing KI and labeled "UNKNOWN C"
- 12 dropper bottles containing AgNO₃ and labeled "UNKNOWN D"
- 12 dropper bottles containing Na₂CO₃ and labeled "UNKNOWN E"
- 12 dropper bottles containing NaCl and labeled "UNKNOWN F"

Other Equipment

- Large well plates